

**SECTION-A**

Note:

- 1) Attempting all MCQs is compulsory. This paper along with the OMR sheet must be returned to the superintendent after due time.
- 2) Fill the circle (A) (B) (C) (D), which one is correct with blue or black ball point in separate OMR Sheet like ●
- 3) If more than one circle in the OMR sheet is filled then no credit will be given to such answer.

I.i. Brass and bronze are the example of:

- (A) Compound      ● Mixture      (C) Element      (D) Substance

ii. Cl is considered as free radical because it has:

- Unpaired Electron      (B) Paired Electron      (C) No Electons      (D) Even no. of electrons

iii. Gold foil was surrounded by photographic plate or:

- (A) ZnO fluorescent screen      ● ZnS fluorescent screen      (C)  $ZnCl_2$  fluorescent screen      (D) Zn Br<sub>2</sub> fluorescent screen

iv. Element "X" is located in period III, group II of periodic table, deduce its electronic configuration.

- (A) 2,8,3      ● 2,8,2      (C) 3,8,3      (D) None of these

v. If electronegativity difference is greater than 1.7, then bond will be:

- Ionic Bond      (B) Covalent Bond      (C) Coordinate Covalent Bond      (D) Metallic Bond

vi. Attraction between highly electron deficient hydrogen and highly electronegative forms:

- (A) Covalent Bond      (B) Ionic Bond      ● Hydrogen Bond      (D) Metallic Bond

vii. Gases molecules have large empty spaces, so they are highly:

- (A) Compressible      ● Diffusible      (C) Compatible      (D) Reliable

viii. What amount of NaOH is required to prepare 0.5M solution in 1 litre:

- (A) 80g      (B) 40g      (C) 30g      ● 20g

ix. Colloids can:

- Scatter light      (B) Not Scatter light      (C) Absorb heat      (D) Evolve heat

x. Which of the following is not correct?

- (A) Oxygen is neutral to litmus      ● Oxygen is not combustible      (C) Oxygen supports combustion      (D) Oxygen is a powerful oxidizing agent

xi. What is the oxidation number of Sulphur in H<sub>2</sub>SO<sub>4</sub>?

- +6      (B) +5      (C) -5      (D) 0

xii. The elements having lower ionization energies have higher tendency.

- (A) To gain electrons      (B) To retain electrons      ● To lose electrons      (D) To share electrons