

RAWALPINDI BOARD

Chemistry (New Scheme)

(Session: 2023)

(Group-I / A-2023)

SSC (Part-I)

Class-9th

New Scheme

Time: 1:45 Hours

(SUBJECTIVE)

Marks: 48

Note: Section I is compulsory. Attempt any Three questions from Section II and any Two parts from Section III.

SECTION-I

2. Write short answers of any Five parts. (5x2 = 10)

- (i) What is radical? Give one example.
- (ii) Differentiate between physical and chemical properties.
- (iii) When does an electron emit or absorb energy?
- (iv) Define electronic configuration.
- (v) What do you know about Dobereiner's triads?
- (vi) Why atomic radius increases down the group in the periodic table?
- (vii) Define ionization energy.
- (viii) Write any two salient features of long form of periodic table.

3. Write short answers of any Five parts. (5 x 2 = 10)

- (i) Point out the type of Covalent Bond in the following molecules. CH_4 , C_2H_6
- (ii) Why HCl has dipole-dipole forces of attraction?
- (iii) Why does ice float over water?
- (iv) Define Evaporation.
- (v) Define Freezing Point.
- (vi) Define Volume/Volume ($\% \frac{V}{V}$) percentage unit.
- (vii) Write two properties of solution.
- (viii) What is Tyndall Effect?

4. Write short answers of any FIVE parts. (5x2 = 10)

- (i) Define oxidation number. Give one example.
- (ii) Find oxidation number of nitrogen in HNO_3 , when oxidation number of H = +1 and O = -2
- (iii) Write function of salt bridge.
- (iv) Define strong electrolytes.
- (v) Write names of two least reactive metals.
- (vi) Define electropositive Characters.
- (vii) Write two uses of Magnesium.
- (viii) Write reaction between Chlorine with hot concentrated NaOH.

SUBJECTIVE PART-II

Questions. Each question has (9 x2 = 18) marks.

- 5.(a) Narrate the properties of molecular ion.
- (b) How many isotopes of hydrogen are there? How they are used in power generation?
- 6.(a) Explain types of Co-Valent bond with one example each.
- (b) State "Charle's Law". Describe its mathematical expression.
- 7.(a) Describe five rules for assigning the oxidation state.
- (b) Explain how are dilute solution prepared from concentrated solutions?