

Model Paper
PR XII
CHEMISTRY (New)
Inter Part-II
(Fresh/Reappear)

Note: Time allowed for Section – B and Section – C is 2 Hours and 40 minutes.

Section – B

Marks: 40

Q-II Answer any TEN parts. Each part carries FOUR marks.

1. Describe the reactions of Phosphorus with Chlorine.
2. Why BeCl_2 is covalent and not ionic?
3. Why transition elements exhibit variable oxidation states? What are their common oxidation states?
4. What is Lassaigne solution? How it is prepared?
5. Discuss the binding energy of transition elements.
6. Define carbocations. Give their types and relative stabilities.
7. How can you distinguish 1 – alkyne from other non-terminal alkynes?
8. Why aliphatic amines are stronger bases than NH_3 ?
9. Describe the Lucas test.
10. Explain briefly glycoside linkage or bond.
11. Explain Electromagnetic spectrum.
12. Give the names of derivatives of carboxylic acids with their functional groups.
13. Why methanoic acid is stronger than ethanoic acid?

Section – C

Marks: 27

Note : Attempt any THREE questions. All questions carry equal marks.

- Q-III** (a) Describe the reactions chlorides of C, Si and Pb with water.
(b) Explain the amphoteric nature of $\text{Be}(\text{OH})_2$
- Q-IV** (a) Give two methods for the preparation of alcohols.
(b) Discuss the effect of substituent on the reactivity of benzene ring.
- Q-V** (a) Explain ozonolysis of alkenes.
(b) Differentiate between the stretching and bending vibrations. What are their various types?
- Q-VI** (a) How does ethanal react with the given reagents?
i. Ethyl magnesium iodide ii. Zn – Hg amalgam and HCl
ii. Lithium aluminium hydride iv. Acidified $\text{K}_2\text{Cr}_2\text{O}_7$
(b) Give IUPAC names to the following compounds.

