

FAISALABAD BOARD

Chemistry (New Scheme)

(Session: 2023)

(Group-I / A-2023)

SSC (Part-I)

Class-9th

New Scheme

Time: 1:45 Hours

(SUBJECTIVE)

Marks: 48

Note: Section I is compulsory. Attempt any Three questions from Section II and any Two parts from Section III.

SECTION-I

2. Write short answers of any Five parts. (5x2 = 10)

- (i) What is relative atomic mass?
- (ii) Define empirical formula with an example.
- (iii) Write the names of two isotopes of hydrogen with their symbols.
- (iv) For what purpose, uranium-235 is used?
- (v) Define electronegativity.
- (vi) How many elements are present in 4th and 6th period?
- (vii) Why the size of atom increases from top to bottom in a group?
- (viii) Define Mendeleev's periodic law.

3. Write short answers of any Five parts. (5 x 2 = 10)

- (i) State octet and duplet rule with an example.
- (ii) Define intermolecular forces.
- (iii) Differentiate between polar and non-polar covalent compounds.
- (iv) Define standard atmospheric pressure and explain with an example.
- (v) What do you mean by Pascal law? How many Pascal's are equal to 1 atm?
- (vi) Differentiate between crystalline and amorphous solids.
- (vii) Why is iodine soluble in CCl_4 and not in water?
- (viii) Define solution. Give an example.

4. Write short answers of any FIVE parts. (5x2 = 10)

- (i) Find the oxidation number of nitrogen in HNO_3 when the oxidation numbers of H = +1 and O = -2.
- (ii) What is meant by reducing agent?
- (iii) What are electrolytes? Write an example.
- (iv) What is the advantage of galvanizing?
- (v) Write any two chemical properties of metals.
- (vi) Write any two uses of magnesium.
- (vii) Write any two uses of gold.
- (viii) Why is HF a weak acid?

SUBJECTIVE PART-II

Questions. Each question has (9 x 2 = 18) marks.

- 5.(a) Define chemical formula. How a chemical formula is written?
- (b) Write four properties of canal rays (proton).
- 6.(a) Explain five properties of metal.
- (b) Write the experimental verification of Boyle's law.
- 7.(a) Define electroplating. Explain electroplating of chromium.
- (b) Explain solubility and solute-solvent interaction along with example.